

5.1.2 How far?

Equilibrium					
(a) use of the terms <i>mole fraction</i> and <i>partial pressure</i>					
(b) calculation of quantities present at equilibrium, given appropriate data					
(c) the techniques and procedures used to determine quantities present at equilibrium					
(d) expressions for K_c and K_p for homogeneous and heterogeneous equilibria					
(e) calculations of K_c and K_p , or related quantities, including determination of units					
(f)					
(i) the qualitative effect on equilibrium constants of changing temperature for exothermic and endothermic reactions					
(ii) the constancy of equilibrium constants with changes in concentration, pressure or in the presence of a catalyst					
(g) explanation of how an equilibrium constant controls the position of equilibrium on changing concentration, pressure and temperature					
(h) application of the above principles in 5.1.2 How far? for K_c , K_p to other equilibrium constants, where appropriate					